

X-RAY STRUCTURE ANALYSIS OF ALL-CIS-1,6-DICHLOROCYCLODECA-1,3,6,8-TETRAENE

Olga Kennard*, D. G. Watson, J. K. Fawcett and K. Ann Kerr

University Chemical Laboratory,

Cambridge,

England.

(Received 25 May 1967)

The stereochemistry, molecular dimensions and crystal structure of 1,6-dichlorocyclodeca-1,3,6,8-tetraene were determined by a three-dimensional X-ray diffraction analysis.

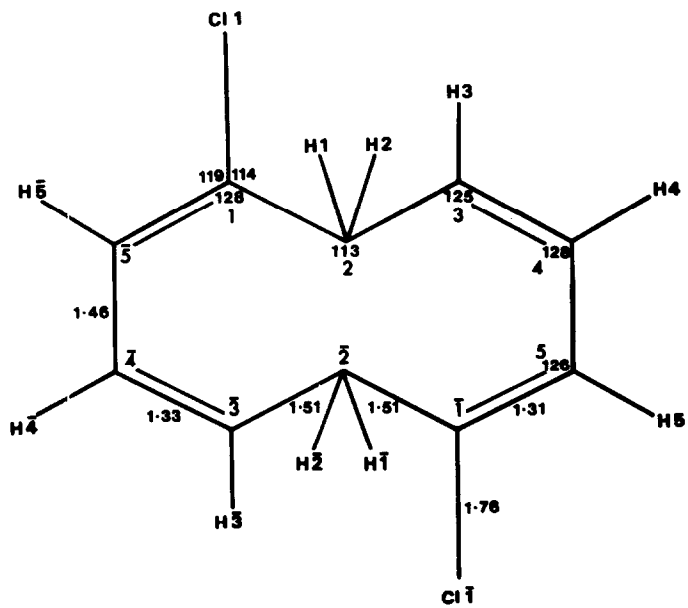
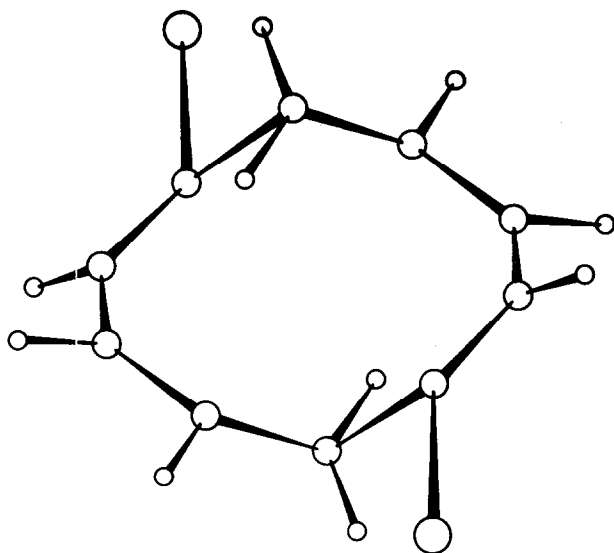
The material was prepared by Professor Sondheimer and Dr. Grohmann as described in the previous communication and the crystallographic analysis undertaken to complement the chemical investigations. The structure was solved directly from the X-ray data, with minimal chemical assumptions. The molecule proved to be centrosymmetric, with an all-cis configuration (Fig. 1.) and with dimensions indicated in Table 1.

TABLE 1.

$\text{Cl}(1) - \text{C}(1) = 1.76\text{\AA}$	$\text{C}(5) - \text{C}(1) - \text{Cl}(1) = 119^\circ$
	$\text{C}(2) - \text{C}(1) - \text{Cl}(1) = 114^\circ$
$\text{C}(1) - \text{C}(2) = 1.51\text{\AA}$	
$\text{C}(2) - \text{C}(3) = 1.51\text{\AA}$	$\text{C}(5) - \text{C}(1) - \text{C}(2) = 128^\circ$
	$\text{C}(1) - \text{C}(2) - \text{C}(3) = 113^\circ$
$\text{C}(3) - \text{C}(4) = 1.33\text{\AA}$	$\text{C}(2) - \text{C}(3) - \text{C}(4) = 125^\circ$
$\text{C}(5) - \text{C}(\bar{1}) = 1.31\text{\AA}$	$\text{C}(3) - \text{C}(4) - \text{C}(5) = 128^\circ$
	$\text{C}(4) - \text{C}(5) - \text{C}(\bar{1}) = 126^\circ$
$\text{C}(4) - \text{C}(5) = 1.46\text{\AA}$	
$\text{C}(2) - \text{H}(1) = 1.0\text{\AA}$	$\text{C} = \text{C} - \text{H} (\text{av.}) = 118^\circ$
$\text{C}(2) - \text{H}(2) = 1.1\text{\AA}$	$\text{C} - \text{C} - \text{H} (\text{av.}) = 112^\circ$
$\text{C}(3) - \text{H}(3) = 0.9\text{\AA}$	$\text{H}(1) - \text{C}(2) - \text{H}(2) = 109^\circ$
$\text{C}(4) - \text{H}(4) = 1.0\text{\AA}$	
$\text{C}(5) - \text{H}(5) = 1.0\text{\AA}$	

*External Staff, Medical Research Council

FIGURE 1.



Each half of the molecule may be defined by two planes; one through atoms C(2), C(3), C(4), C(5), H(3), H(4) and one through atoms C(4), C(5), C($\bar{1}$), C($\bar{2}$), H(5), Cl($\bar{1}$). The two planes are inclined at an angle of 57.7° to each other. This arrangement of atoms allows for a separation of 2.5\AA between the two inner hydrogens H(1) and H($\bar{1}$). The bond lengths and angles found agree closely with expected values.

Crystal Data: $C_{10}H_{10}Cl_2$. M. Wt. = 201.1. Chunky prisms from petroleum ether (b.p. $40-60^\circ$). Diagonal extinction on the prominent, diamond-shaped faces.

Monoclinic: $a = 7.577\text{\AA}$, $b = 7.878\text{\AA}$, $c = 8.397\text{\AA}$, $\text{av. } \sigma = 0.001\text{\AA}$, $\beta = 109.1^\circ$
 $Z = 2$ $D(\text{calc}) = 1.43 \text{ g/cm}^3$. Space group $P2_1/c$ from absences. Molecular symmetry: $\bar{1}$.

Experimental: The crystals were unstable and had to be enclosed in capillary tubes, filled with nitrogen. After preliminary X-ray photographs, all measurements were carried out on a Picker 4-circle X-ray diffractometer. Of the 863 points measured to a 2θ maximum of 136.5° , 838 were of significant intensity.

Solution and Refinement: The position of the chlorine atom of the asymmetric unit was deduced from a three-dimensional sharpened Patterson map and the choice between alternative locations was made on the basis of the intermolecular chlorine approaches. The coordinates of the carbon atoms were obtained from Fourier maps and the structure refined through three cycles each of isotropic and anisotropic least squares calculations using the X-ray 63 system of programs. The hydrogen contributions were next inserted and a further four cycles of mixed anisotropic/isotropic refinement with 70 variables led to a reliability index of 9.0% (observed terms only). The average values of the standard deviation in atomic coordinates are 0.002\AA for Cl, 0.006\AA for C and 0.07\AA for H. Refinement is continuing.

Acknowledgements: We would like to thank Professor Lord Todd for laboratory facilities, the Science Research Council for the purchase of the diffractometer, the Medical Research Council and the Office for Scientific and Technical Information for their support. We would also like to acknowledge some very helpful discussions with Dr. C. Romers.